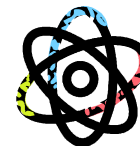


Agenda for Week of: Dec 13, 2010

Classroom Character Theme: _____

Chapter 4 Sections 1 and 2 Focus Questions (Thurs & Fri):

- 1) What are "Atoms" made of and why don't they ever "touch"?
- 2) How can matter (all "stuff") be made of ...empty space?!?



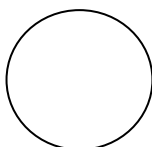
Date	Objectives	Vocabulary	Homework
Monday 12/13	Organize notebooks & review for tomorrow's test	Review all vocab from previous Chapter 3 agendas	1) Complete review sheet 2) Staple your Unit 2 packet together 3) Have progress report signed and returned by Wednesday, 12/15
Tues 12/14	Chapter 3 TEST		<i>Extra credit:</i> Finish the puzzle from after the test
Wed 12/15	Empower3000		<i>Extra Credit:</i> Read Chapter 4 Section 1 and answer p. 118 #1-4
Thurs, 12/16	List the parts of an atom and their relative size and charge	Atom Proton, neutron electron	1) Read Ch 4.2 (pages 119 – 127) 2) List and define the highlighted vocab 3) Make a TABLE to answer p. 127 #1
Fri, 12/17	Contrast atomic # & mass #	atomic number mass number	None

REMINDERS

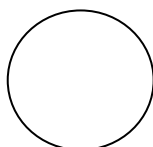
- 1) **Calculator:** Remember to bring your calculator to class every day!
- 2) **Binder:** Bring it to class **every day**. You will have a weekly Binder Check Pop-quiz.
- 3) **Homework** counts for 10% of your grade and is supposed to HELP your grade!
- 4) Students who fail to complete homework may receive **academic detentions**.
- 5) **Have questions?** Stay for **tutoring** on Mondays and Thursdays HERE from 2:45 – 3:30. You will earn **extra credit** for showing this extra effort!

Review Sheet for Chapter 3 Test

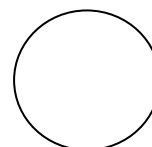
- 1) List the three states of matter (in addition to "plasma"): _____, _____, and _____
- 2) Draw what *molecules* look like in these states of matter. This represents a _____scopic model



Solid



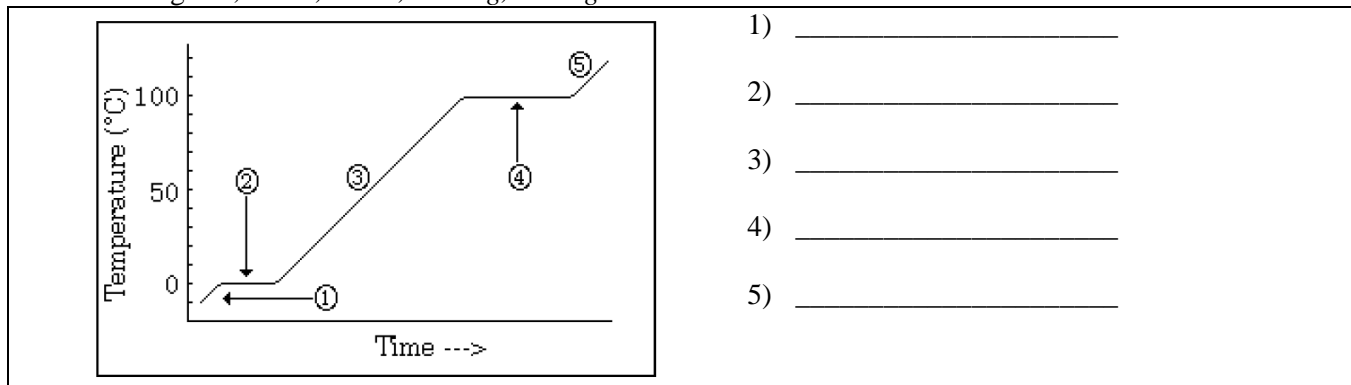
Liquid



Gas

- 3) These two states of matter CANNOT be compressed: _____ and _____
- 4) This state of matter has a fixed volume but NOT a fixed shape: _____
- 5) This state of matter has atoms that are the MOST tightly held together (as though they had invisible springs connecting them): _____
- 6) "Changes of state" are a *chemical* / *physical* property and include sublimation, deposition, _____, _____, _____
- 7) During melting, energy is *absorbed by* // *released from* the molecules, which causes them to move *faster* / *slower*.

- 8) Label each part of the graph below by identifying what is happening on the line to the right. Use the following: *ice, water, steam, melting, boiling*



- 9) During 'changes of state,' the temperature *does / does not* rise because the heat energy is turned into _____ energy and causes the molecule's motion to *increase / decrease*.

- 10) List **two** examples of UNITS used for each of these variables:

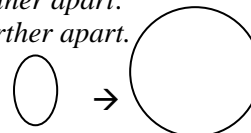
Pressure: _____, _____ Temperature: _____, _____ Volume: _____, _____

- 11) According to _____ Law, the VOLUME of a gas will *increase / decrease* as the PRESSURE increases (assuming the temperature stays the same).

- 12) When a gas inside a balloon cools, **the molecules** will move *closer together / farther apart*.

- 13) When a gas inside a balloon warms, **the molecules** will move *closer together / farther apart*.

- 14) According to _____ Law, the balloon to the right got *hotter / colder*.



- 15) According to _____ Law, the TEMPERATURE of a gas will *increase / decrease* as the PRESSURE increases (assuming the volume stays the same).

Boyle's Law: $V_2 =$ $P_2 =$

- 16) A balloon has a volume of 500 mL at a pressure of 1.0 atm. It is taken under water to a depth where pressure increases to 2 atm. What is the new volume of the balloon?

Variables Equation Work (numbers in the formula)

$P_1 =$

$V_1 =$

$P_2 =$

$V_2 =$

Final answer with units:

- 17) A sample of gas in a syringe has a volume of 19 mL at a pressure of 48.7 kPa. The plunger is depressed until the volume reaches 13 mL. What is the new volume?

Variables Equation Work (numbers in the formula)

$P_1 =$

$V_1 =$

$P_2 =$

$V_2 =$

Final answer with units: